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AN ELECTRON SPECTROSCOPY STUDY OF AMMONIA ADSORPTION ON CLEAN A--ETC(U)

FEB 80 J W ROGERS, C T CAMPBELL, R L HANCE

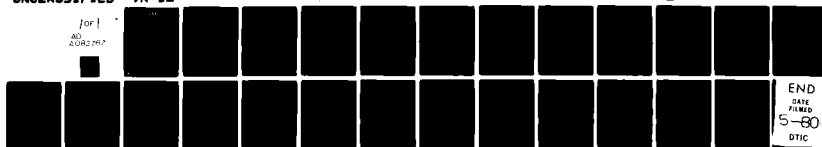
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⁶ An Electron Spectroscopy Study of Ammonia

Adsorption on Clean and Oxidized Aluminum.

by

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**An Electron Spectroscopy Study of Ammonia
Adsorption on Clean and Oxidized Aluminum^a**

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Abstract

Photoelectron spectra for ammonia adsorbed in submonolayer and multilayer amounts on clean and oxidized aluminum have been measured and interpreted. Uptake at 128K is dominated by weak molecular adsorption and saturates at submonolayer amounts whereas at 106K multilayers can be formed. A tiny amount of dissociative adsorption is noted. Dissociation is promoted by exposure to Ar⁺ ion bombardment. At 128K the initial sticking coefficient is 0.13 ± 0.08. On oxidized Al, as compared to clean, the saturation amount of adsorbed NH₃ is larger and it is more tightly held. In both cases the work function change is consistent with a model in which the ammonia adsorbs with the nitrogen atom toward the surface. On clean Al, there is good evidence that NH₃ tends to be physisorbed with the major bonding arising between the dipole of NH₃, and the image dipole induced in the metal. This is supported by the resemblance of the separation of the valence orbitals to the separation in the gas phase, as well as their shift to higher binding energy with increasing coverage. On oxidized Al the bonding is similar but the bond strength is greater. There is no evidence for significant NH₄⁺ formation on either clean or oxidized Al. The Lewis acidity of the substrate increases with the extent of surface oxidation. The bonding of NH₃ to oxidized Al is very similar to the intermolecular bonding in an ammonia molecular crystal.

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1. Introduction.

Alumina is an important catalyst support material and has intrinsic catalytic activity for certain reactions[1]. Being an ionic solid it has acid/base sites (Lewis and Brönsted-Lowry) associated with its surface cations and anions. The nature and behavior of these sites are of fundamental importance both in the presence and absence of dispersed metal particles.

Characterization of acidic sites on metal oxides has been studied using several techniques but direct spectroscopic data on the nature of such sites is scarce and much of our knowledge comes from inferences about the reactions that occur on them. Studies of bonding of gaseous bases to solid acid sites have been carried out using IR techniques [2,3,4,5,6]. Surface protonic sites, as well as the adsorption of deuterated pyridine on Brönsted sites, have been studied using NMR techniques [7,8,9]. Other techniques are discussed in reviews by Forni and Goldstein [10,11]. These techniques suffer from the two major disadvantages that the surfaces are usually not clean and well defined and that high pressures of the adsorbate gases are necessary to produce acceptable signal levels. In addition most of the techniques fail to differentiate between Lewis and Brönsted type acid sites, that is, they measure only total surface acidity.

XPS and UPS techniques, which probe electron levels within the adsorbate as well as bonds between the surface/adsorbate complex, should be ideal for the study of these interactions. Electron spectroscopic techniques have received little attention in examining surface acid sites because of experimental difficulties inherent with working on insulators. We have recently shown that charging

and thermal conductivity problems can be overcome by working with well characterized, selectively oxidized metal surface [12]. In addition, Nannava *et al.* have shown that oxidized Al surfaces model bulk γ - Al_2O_3 both physically and chemically [13].

Here we present an investigation of the interaction of NH_3 with clean and partially oxidized Al surfaces at low temperature using XPS and UPS techniques. XAES results from the same study have recently been published elsewhere [12] and the results are summarized here. In addition a characterization of the partially oxidized Al surfaces is presented. As inferred above, the motivation for this study comes from our desire to understand the molecular level details of the atomic and electronic structure of species chemisorbed on well-characterized metal oxides.

To contrast the differences in the behavior of NH_3 on clean and oxidized Al surfaces, it was first necessary to investigate the NH_3 interaction on clean Al. Electron spectroscopic studies of NH_3 adsorption have been reported on several metals including Fe [14,15,16,17], Cu[18], W[19,20], Ru[21,22], Mo[23,24] and Pd[25] as well as oxidized Cu[18]. Investigations of the interaction of NH_3 with metal oxides using these techniques are scarce [26].

2. Experimental Techniques.

The substrate was 1 cm^2 of 99.999% Al foil mounted such that it could be cooled to 106 K. The surface temperature was monitored with an attached chromel-alumel thermocouple.

The Al was initially difficult to clean; all of Ar ion bombardment at $0.8 \text{ milliwatt/cm}^2$ (5KV) at room temperature was required to remove the oxide layers as judged by XPS. Subsequent oxide layers acquired by overnight adsorption of residual gases or by O_2 exposure were removed with ≈ 30 minutes of sputtering under the above conditions. Throughout this paper the Al will be called "clean" if the $\text{O}(1s)$ integrated peak area was less than 5% of that observed for a monolayer. A monolayer, using the convention of Martinson *et al.* [27] is defined as the exposure necessary to produce an $\text{O}(1s)$ peak area 90% of that at saturation. At 128 K this corresponds to an O_2 exposure of $\approx 8 \text{ L}$. Freshly sputtered Al was not annealed prior to exposure.

The experiments were performed on a PHI Model 548 Electron Spectrometer. The electron energy analyzer was operated in the retarding mode at a constant resolution of 0.4 eV (FWHM) for UPS and 1.6 eV for XPS. The surface was biased with a small negative voltage (3-6 volts) to facilitate accurate measurements of the kinetic energy distribution widths which were used to calculate the photoelectric work function [28]. All binding energies

were referenced to the Fermi level of Al; the $\text{Al}(2s)$ BE for the clean surface was 117.6 eV, in good agreement with reported values [29].

The He resonance lamp, used for production of He $\text{I}(\epsilon=21.2 \text{ eV})$ and He $\text{II}(\epsilon=40.8 \text{ eV})$ photons, was differentially pumped.

The system base pressure, 2×10^{-10} Torr, rose to 1.2×10^{-8} Torr of 99% He when the line-of-sight valve into the UHV was open for UPS measurements. $\text{Mg(K}\alpha)$ x-rays ($\epsilon=1253.6 \text{ eV}$) were used for XPS measurements.

NH_3 exposures were accomplished using a carefully calibrated, dynamically pumped doser system equipped with a multi-channel array which helped eliminate flux gradients across the sample [12]. An exposure of 50 L ($1 \text{ L} = 1 \text{ Langmuir} = 10^{-6} \text{ Torr-sec}$) of ultra-pure NH_3 could be achieved in 13 minutes without raising the UHV system pressure above 1×10^{-9} Torr. Oxygen exposures were accomplished by backfilling the UHV system to no more than 2×10^{-7} Torr for a set amount of time. Residual and desorbed gases were monitored with a quadrupole mass spectrometer.

The XPS and UPS data were taken digitally using signal averaged pulse counting techniques and were stored in a multi-channel analyzer. A 20 eV wide energy window was typically scanned 16 times at a rate of 50 msec/channel and stored in 256 channels of memory. These could be permanently transferred to a magnetic tape on a CDC 6600 computer for subsequent manipulation.

3. Results.

3.1 The Characterization of Partially Oxidized Al Surfaces.

The O(1s) transition is plotted in Fig. 1 for several O₂ exposures at 128 K. The residual oxygen signal on "clean" Al always represented <5% of one monolayer of oxygen.

As noted in the UPS results below, this residual oxygen is probably dispersed in the bulk and therefore amounts to a small surface impurity. With increasing O₂ exposure a symmetric O(1s) peak is observed, the area of which grows until saturation after 150 L. The peak energy (531.3 eV) and FWHM (3.0 eV) do not change with exposure.

The O(1s) satellite

from Mg K $\alpha_{3,4}$ radiation, clearly present at higher exposures, is situated at ~522 eV.

The FWHM of the Al(2s) peak (not shown) grows from 2.5 eV on clean Al to 3.0 eV after 9 L exposure, and thereafter it remains constant. This is attributable to a shift of some intensity from 117.6 eV (Al⁰) toward 120.2 eV (Al³⁺) as the surface oxidizes [30].

Interestingly, condensing several multilayers of NH₃ on the oxidized surface caused the 2s intensity from oxidized Al at 120.2 eV to increase relative to the clean Al(2s) intensity at 117.6 eV, because the Al signal was coming from nearer the surface as the depth of NH₃ increased. This suggests that the oxygen resides primarily in the topmost several layers, which agrees with the general picture of the oxidation process [31,32,33]. The high signal

level, the instability of O₂ multilayer formation at these temperatures [34] and the lack of any attenuation of the Al(2s) transition, indicate that an appreciable amount of oxygen is being incorporated into the sub-surface region as previously reported [35,36].

Both the He I and He II UPS regions for the uptake of oxygen at 128 K are shown in Fig. 2. Curves a ($\hbar\omega=21.2$ eV) and e ($\hbar\omega=40.8$ eV) show the valence band of Al after sputter cleaning. The Fermi level is clearly present in both cases and BE's are referenced to it. The clean Al spectrum often showed a broad, low intensity peak centered near 7.1 eV due to emission from the O(2p) resonance (see Fig. 6). This peak derives from a small and variable amount of residual oxygen (~5% of a monolayer) which is detected with high sensitivity by UPS. The adsorption coefficient for 40 eV photons is twenty times greater for Al₂O₃ than Al [37]. Hagstrom, et al. have shown that ~0.01 monolayers of surface oxygen on Al can be detected by UPS [38].

The peak area in Fig. 2 grows rapidly with increasing exposure. After an exposure of 4.3 L, the Fermi level is no longer distinguishable because oxygen ions are formed and remove electron density from the valence band and because satellite peaks in He II add intensity just to the right of E_F in Fig. 2. The asymmetry of the O(2p) peak (FWHM = 6 eV) is discussed below. Curve i shows a low intensity transition near zero BE which is O(2p) emission caused by a He II satellite ($\hbar\omega = 48.31$ eV).

$\sim 4 \text{ \AA}$ [40,41], we suggest that the 9.3 eV peak is characteristic of oxygen below the surface. This is to be compared with other assessments of the O(2p) splitting [42,43,44].

The integrated peak areas from the O(1s) and O(2p) ($\sim 40.8 \text{ eV}$) regions as a function of oxygen exposure (at 128 K) are shown in Fig. 3. The coverage quickly rises to half its maximum value at an exposure of 2.5 L and continues to increase monotonically until it saturates at $\sim 30 \text{ L}$. Using a first order Langmuir least-squares fit to the data and assuming saturation coverage = $5.2 \times 10^{14} \text{ cm}^{-2}$, an initial sticking probability $S_0 = 0.12$ was obtained.

Our purpose was to characterize the oxidized surfaces such that oxide-related intrinsic effects could be separated from the adsorbate-induced extrinsic effects accompanying NH_3 adsorption. The characterization was comprehensive enough to insure that the present results were in agreement with those previously reported in the literature. From the great body of information on Al oxidation the following picture of partially oxidized Al surfaces emerges:

1. At 0.9 L exposure, oxygen penetrates below the first layer of Al, producing a depletion layer rich in Al^{3+} as compared to clean Al [45,46].
2. Intermediate exposures of O_2 (4.3 L) produce a surface $\sim 75\%$ covered by amorphous Al_2O_3 islands of thickness no greater than $\sim 7 \text{ \AA}$ [30,33,36,38]. Patches of Al and Al^{3+} probably exist on this surface.

The He I region, shown at the left of the figure, is plagued by a high secondary electron background at low KE which is caused in part by the analyzer transmission varying as KE^{-1} in the retarding mode. On the cleanest Al surfaces, a narrow, low intensity peak is present at $\sim 9.3 \text{ eV}$ as seen in a. With increasing exposure, the intensity at 7.2 eV in He I grows faster than the 9.3 eV peak and becomes comparable to the intensity of the latter by $\sim 1 \text{ L}$. After an exposure of 4.3 L the O(2p) resonances in He I and He II (spectra d and h) are indistinguishable in general peak shape.

The width and asymmetry of the O(2p) peak has been explained in terms of a superposition of two narrow transitions separated by 2.3 eV [30]. This interpretation suggests that O(2p) orbitals normal and parallel to the substrate have different SE for oxygen atoms in equivalent positions on the surface. Our results suggest that these two peaks are clearly resolvable in He I but not in He II because of an enhancement in the relative intensity of the 9.3 eV peak in He I. No more than 18% of this enhancement could be due to analyzer transmission. Since the electron escape depth is much larger in He I than in He II (20-30 \AA [39] versus

3. Heavy oxidation (260 L) leads to a 4-7 Å amorphous Al_2O_3 film covering the entire surface [30]. The observed behavior of NH_3 adsorbed on these oxidized surfaces is consistent with this picture.

3.2 The Interaction of NH_3 With Clean Al.

The N(1s) region for the adsorption of NH_3 on clean Al at 128 K is shown in curves a-e of Fig. 4. The spectra all exhibit a major peak near 400 eV, the area of which increases monotonically with increasing exposure. This is near the BE at which $\text{NH}_3(\text{ads})$ is expected based on NH_3 adsorption on other metals [14,15,16]. The FWHM of 3.0 eV remains constant as the exposure is increased, but the BE increases from 400.0 to 400.4 eV at exposures greater than 38 L which is close to saturation at 128 K.

A broad, low intensity peak at 396.2 eV is present even at the lowest exposures. The peak area, though difficult to determine accurately, does not appear to change with increasing NH_3 exposure. A ghost peak, originating from $\text{Cu}(\text{L})$ excitation of the Al(2p) transition, was only partially responsible for this feature.

A brief (90 sec.) exposure of the surface which produced curve e (maximum NH_3 adsorption at 128 K) to an Ar^+ ion beam causes a marked decrease in the intensity of the transition near 400 eV while increasing the intensity at 396 eV. This indicates the latter is produced by a dissociated nitrogen-containing species $\text{NH}_{3-x}(\text{x} = 1-3)$, probably $\text{N}(\text{ads})$. A $\text{N}(\text{ads})$ species with a BE at 397 eV has been reported on Al[32] and at 397 ± 0.5 eV on W and Fe[14,15,16,47] which lends further evidence to the present assignment. There is no evidence for NH_4^+ which would give a peak on the high BE side of the NH_3 signal.

Multilayer formation is not possible at 128 K and curve

e represents saturation coverage at this temperature. Spectrum f shows an NH_3 multilayer ($58 \text{ L} \approx 5$ molecular layers) stabilised by lowering the surface temperature to 106 K. The width of the NH_3 multilayer peak in f (FWHM = 2.4 eV) is much narrower than that of the adsorbed NH_3 layer (FWHM = 3.0 eV) shown in e. The 0.6 eV broadening of adsorbate core levels with respect to the multilayer will be discussed below. A satellite peak near 391 eV in spectrum f is due to $\text{MgK}\alpha_{3,4}$ radiation ($h\nu = 1262.9 \text{ eV}$) from the achromaticity of the X-ray anode.

An understanding of the molecular orbital structure and UPS spectrum of gas phase NH_3 is central to the assignment of photoemission peaks derived from adsorption of NH_3 on Al. The orbital representations of the valence levels for NH_3 showing contours of constant charge density [48] are presented in Fig. 5. From these diagrams it is apparent that the electrons in the $3a_1$ orbital are predominantly non-bonding, while the $1e$ orbitals have high electron density centered between the N and H atoms and form the main N-H bonds.

The UPS ($h\nu = 21.2 \text{ eV}$) spectrum of NH_3 gas [49] is also shown in Fig. 5. The bands, whose maxima occur at

10.88 and 16.0 eV, arise from the $3a_1$ and $1e$ orbitals respectively. The $1e^{-1}$ transition is broad (FWHM = 2.2 eV) and exhibits a vibrational envelope. Interestingly, vibrational fine structure also appears in the non-bonding lone-pair $3a_1^{-1}$ band. This is attributed to the NH_3^+ ion being planar as opposed to NH_3 being trigonal bipyramidal; so despite non-bonding character of the $3a_1$ orbital, it is important in

determining the NH_3 geometry [50,51].

The UPS ($h\nu = 40.8 \text{ eV}$) spectra for submonolayer coverages of NH_3 on clean Al at 128 K are displayed in Fig. 6. Curve a, showing the valence band of "clean" Al, is not identical with Fig. 2c, and indicates the variability of residual oxygen which we found. As seen in b-e, upon exposure to NH_3 , two peaks are present centered at 6.3 and 11.8 eV; their peak areas increase with increasing exposure. The intensity at the Al Fermi edge is rapidly masked by a satellite peak produced from He II ($3s + 1s$) ($h\nu = 48.31 \text{ eV}$) excitation of the nitrogen orbitals which overlaps E_F .

The gas phase vertical ionization potentials are shown at the top of Fig. 6 after referencing them to E_F of Al by subtraction of the work function of the Al surface saturated with NH_3 at 128K [52]. On this scale, the $1e^{-1}$ and $3a_1^{-1}$ for gas phase ammonia occur at 13.9 and 8.4 eV, respectively. There is some uncertainty in the vertical ionization potential for the $1e$ orbital because the transition is broad and asymmetric. Here we have followed Grunze et al. [17] and have used the midpoint of the FWHM in Fig. 5. If instead, the typical method of using maxima is used, then the $1e^{-1}$

would be placed at 13.5 eV. We favor the former largely because the latter leads to contradictory conclusions about the charge density on the nitrogen atom in adsorbed ammonia.

Comparing the gas and adsorbed phase spectra, it is clear that the symmetric features with maxima at 11.8 and 6.3 eV in the latter correspond to the gas phase $1e^{-1}$ and $3a_1^{-1}$ transitions respectively. Between 10 and 38 L, both these transitions shift to higher binding energy by about 0.4 eV. With respect to the gas phase, both adsorbed phase transitions are shifted to lower binding energy by 2.1 eV. We must obviously allow some uncertainty in these shifts, particularly in the case of the $1e^{-1}$. Regardless of this uncertainty the splitting of the adsorbed NH_3 on Al (5.5 eV) more like the gas phase than it is on iron (4.4 eV) [17]. This fact suggests that chemical bonding of NH_3 , which on iron is thought to occur mainly through the lone pair of nitrogen, is largely absent in the case of Al. Another factor favoring weak bonding is the uniform shift to higher binding energy (0.4 eV) with coverage. This can be understood in terms of interactions between neighboring dipoles and their images [53]. This model gives a shift to higher binding energy because the permanent dipole of an ammonia molecule is stabilized through interactions with more than one image dipole. This attractive interaction is not fully compensated by repulsive interactions between dipoles. A third observation favoring very weak binding of NH_3 to clean Al comes from warming experiments ($N(1s)$ versus T). Our results indicate a heat of adsorption of 9 ± 2 kcal mole $^{-1}$.

In spite of these considerations, there is one observation that suggests a stronger coupling of the lone pair orbital with the Al

surface than the N-H bonding orbital; namely, the branching ratio. This quantity defined as the ratio of peak areas, $1e^{-1}/3a_1^{-1}$, is 2.0 for the gas [54] and 2.26 for the adsorbed phase [55]. Assuming that the $1e$ orbital is non-bonding and that the cross section for its ionization is comparable to the gas phase value, the $3a_1^{-1}$ intensity is clearly less than its gas phase analog suggesting a surface interaction with the lone pair which effectively removes some electron density from around the nitrogen atom. In the present case, this appears to be a redistribution spatially of the charge in the $3a_1$ orbital rather than an actual loss of electron density. Accordingly, a shift of the $3a_1^{-1}$ binding energy to higher values is expected but not observed. One possible explanation is compensation due to more effective core-hole screening in the $3a_1^{-1}$ orbital due to its closer proximity to the surface. Calculations [56] indicate that differentials of a few tenths of an eV could be due to this source.

Taken together, all this evidence points to a very weak interaction of molecular NH_3 with clean Al surfaces resulting in a relaxation energy shift, ΔE_R , of 2.1 eV based on the position of the $1e^{-1}$ transition. While the $3a_1^{-1}$ shifts by the same amount, its intensity drops from gas phase values (relative to $1e^{-1}$) suggesting a slight redistribution of charge in this orbital as it interacts with the surface. We also find that work function changes (presented below) are consistent with the lone pair end of the NH_3 molecule toward the surface.

Comparison of UPS results for condensed and adsorbed NH_3 is perhaps even more informative than comparison of the latter with the gas phase results. Table I summarizes gas phase [49,54,57].

condensed [58], and the present physisorbed data. In passing from the physisorbed to the condensed phase the peak widths narrow 0.7 eV (FWHM). This is comparable to the 0.6 eV narrowing of the N(1s) core ionization noted above. Increased hole lifetimes in the multilayer can account for only a small part of this increase. For the case of a monolayer and multilayer of CO adsorbed on Cu [59,60] the observed decrease was 0.1 eV FWHM and this is in line with theoretical estimates [61]. The remaining 0.5 eV broadening in adsorbed NH_3 is tentatively ascribed to the range of work functions present on freshly sputtered polycrystalline Al. To the extent that the photoionization peaks should track the local vacuum level for physisorbed species [62,63] these peaks should broaden as much as the range of work functions available. Indeed it has recently been observed that the Xe levels broaden or even split when Xe is physisorbed on a partially reconstructed Ir(100) surface (where the work functions for reconstructed and unreconstructed Ir(100) both exist) [64]. This kind of heterogeneity-related broadening also explains why the NH_3 peaks we observe on Al are about 1 eV broader than those on Fe (111) [17].

From the linewidths and branching ratios presented in Table I, we are led to believe that the bonding of NH_3 in a condensed multilayer and a physisorbed monolayer is very similar. Cohesion in molecular crystals of NH_3 results from dispersion forces and weak hydrogen bonds [65,66]. As mentioned above we estimate from warmup measurements that the heat of adsorption of NH_3 on Al is 9 ± 2 kcal/mole whereas the heat of sublimation for NH_3 is 6.5 kcal/mole [67]. These facts are consistent with the notion that the Al

atoms involved in the NH_3 monolayer are similar to the H atoms in the NH_3 multilayer insofar as binding energetics and molecular orbital structures are concerned.

Spectrum b of Fig. 6 shows some additional intensity in the region between the $\tilde{\nu}^{-1}$ and $3\tilde{\nu}_1^{-1}$ transitions not present in c, d and e. This is the region where features due to dissociated NH_3 have been reported on other metals [17,32] and we tentatively assign this intensity to NH_3 -x (x = 1-3). However, we cannot rule out the possibility that it may be due to a small oxygen impurity.

The relative NH_3 coverage as a function of exposure, estimated by integrating the areas of the N(1s) peaks, is shown in Fig. 7. Also shown are the summed peak areas from the UPS He II region (Fig. 6). There is good agreement between the XPS and UPS data (normalized at 115 L). This is consistent with a relatively weak NH_3 -Al interaction that involves very little charge transfer. In such a case the perturbation of the adsorbate valence levels is small and nearly coverage independent. Consequently the UPS intensity will reflect the adsorbate concentration.

The inset of Fig. 7 shows the work function change, $\Delta\phi$, as a function of NH_3 exposure at 128 K. These changes were determined by measuring the width of the photoelectron kinetic energy distribution. The measured work function of clean Al was $\phi_{\text{Al}} = 4.2$ eV. The results indicate a sharp decrease in work function accompanying increased NH_3 exposure up to 40 L at

which point the limiting change, $\Delta\phi = 1.7$ eV, is reached. The direction of the work function change is consistent with the nitrogen atom (electronegative end of the NH_3 dipole) being oriented toward the surface and the hydrogens toward the vacuum.

The coverage rises rapidly with NH_3 exposure reaching 80% of saturation at an exposure of 12 L and saturating with less than 60 L exposure. The saturation coverage can be estimated in a number of ways. Based on the absolute $\text{N}(1s)/\text{Al}(2s)$ intensity ratio and using a method described in the literature [32,67], the saturation coverage N_{sat} is $8.6 \times 10^{14} \text{ cm}^{-2}$. This estimate is expected to be high since significant $\text{Al}(2s)$ intensity, appearing as intrinsic plasmon loss structure, is neglected. Using the relation for total work function change [68]:

$$\Delta\phi_{\text{sat}} = 4\pi eM N_{\text{sat}}$$

and approximating the dipole moment (M) of the adsorbate/surface complex with that of gaseous ammonia (1.47 Debye), we obtain $N_{\text{sat}} = 3.2 \times 10^{14} \text{ cm}^{-2}$. A close-packed monolayer of NH_3 (assuming $r_{\text{NH}_3} = 1.53 \text{ \AA}$) contains $1.4 \times 10^{15} \text{ cm}^{-2}$. Saturation coverage at 128 K then corresponds to about 25 to 50% of a full monolayer. To estimate the initial sticking probability the data was least-squares fitted (solid line in the figure) to a first-order Langmuir adsorption equation of the form

$$\theta = (1 - \exp -4.9 \times 10^{14} N_{\text{sat}}^{-1} S_0 L) \quad (3)$$

where N_{sat} is the adsorbate surface density at saturation, S_0 is the initial sticking probability, L is the exposure (in Langmuirs)

and θ is the fractional coverage. Using $N_{\text{sat}} = 5 \times 10^{14} \text{ cm}^{-2}$ as an intermediate value, we obtain $S_0 = 0.13$. We must allow a factor of ± 0.08 to account for uncertainties in N_{sat} .

If the surface temperature is lowered to 106 K, NH_3 condensation is possible. Fig. 8 contrasts the adsorption of NH_3 at 106 K (triangles) and 128 K (circles) from results in Fig. 7). The $\text{N}(1s)$ peak area is plotted as a function of NH_3 exposure and the areas are normalized to the area of the $\text{Al}(2s)$ transition obtained for clean Al prior to each NH_3 exposure. From measurements of the attenuation of the $\text{Al}(2s)$ -signal the multilayer depth can be estimated at several coverages in Fig. 8. Assuming the thickness of an adsorbed NH_3 monolayer to be 3.05 \AA (calculated from van der Waal's radii) and an $\text{Al}(2s)$ attenuation length of 17 \AA [39], the first monolayer occurs at a relative coverage of 0.94. At 128 K, saturation coverage occurs near 0.3, indicating that only 30% of the amount of NH_3 present in a monolayer at 106 K is stable on the surface at 128 K. This is in reasonable agreement with the estimates made above. The remaining data in Fig. 8 will be discussed below.

The uniform 0.4 eV binding energy shift with coverage of the $\text{N}(1s)$ core and $1e$ and $3a_1$ valence levels is noteworthy. As mentioned earlier, this can be understood in terms of long range image charge overlap [53]. Another possibility involves the question of whether adsorbate levels track the vacuum or Fermi level. This has been discussed by Kuppers et al. [64,69] who find similar effects in spectra of Xe on $\text{Pd}(110)$. Physisorption on MoS_2 also exhibits these effects [62,63]. In weak adsorption systems adsorbate levels

are not in full electrochemical equilibrium with the substrate. As a result they do not precisely track the Fermi edge as is the case for chemisorbed systems. Rather they tend to track the vacuum level which changes with adsorbate induced work function changes. The NH_3/Al case is somewhere between the very weak physisorbed and fully chemisorbed cases; therefore, small shifts (≈ 0.4 eV) with much larger changes in work function (1.7 eV) can be seen.

3.3. The Interaction of NH_3 on Partially Oxidized Al Surfaces.

The adsorption of NH_3 on Al surfaces pre-oxidized with 0.9, 4.3 and 260 L O_2 was studied at 128 K. These exposures represent low, half and saturation oxygen coverages at this temperature. The UPS and work function changes accompanying NH_3 adsorption as well as the temperature dependence of the adsorbate-derived XPS features indicate that, just as for clean Al, NH_3 is weakly held on oxidized surfaces; however, the interactions do differ significantly. The N(1s) core level transition appears at 400.3 eV as expected for molecular NH_3 adsorption and is equal to the BE observed for saturated NH_3 on clean Al at 128 K. As shown by the filled squares of Fig. 8 the capacity of the surface for adsorbing NH_3 increases with O_2 predose.

The UPS ($\bar{\nu}_w = 40.8$ eV) features for the adsorption of NH_3 on oxidized Al surfaces are displayed in Fig. 9. As compared to clean Al, there are changes in the BE and intensities of the bands.

To determine if the UPS features of NH_3 adsorbed on oxidized surfaces were simply superpositions of NH_3 and O_2 interactions on clean Al, the appropriate individual NH_3 and O_2 spectra were added, with scaling, to give good visual fits to the results obtained for NH_3 adsorbed on oxidized Al. The following procedure was used. The digitized spectra for ammonia and for oxygen were (1) aligned at the He I Fermi level (not shown), (2) smoothed once with a five point quadratic/cubic smoothing algorithm, [70] (3) rescaled, (4) arithmetically summed, and (5) compared with the actual data from the oxidized surface. The scaling coefficients were varied and the procedure repeated until the best fit to the data (determined by inspection) was obtained.

The dotted curves in the figure show the scaled $O(2p)$ intensity on clean Al at the indicated exposures and the dashed curves show the scaled NH_3 valence band for a 57 L exposure. The superpositions of these spectra are depicted by the dot-dashed curves and the solid lines show the actual data for NH_3 adsorption on oxidized Al. From the overall similarity of these, it is obvious that NH_3 exposures to oxidized Al at 128 K also produce molecularly adsorbed NH_3 . This is in agreement with our XAES results [12] and XPS results discussed below. We will now discuss the slight differences that are observed in these two curves.

The scaling parameters, which

accompany panels (A), (B) and (C) of Fig. 9 are summarized in Table II. Since the NH_3 exposure was fixed (58 L) the coefficient of the NH_3 dose indicates the amount adsorbed under each oxygen pre-dosed condition relative to clean Al. In each case the coefficient is greater than unity implying more NH_3 is adsorbed on oxidized Al. The XPS peak areas are in agreement with this observation as shown by the squares in Fig. 8. These results are all consistent with an increasing Lewis acidity with the extent of surface oxidation. An increased surface area would also tend to increase the amount of adsorbed NH_3 , but photoelectron spectra and thermal stability experiments described below indicate that more than an area change is involved.

The oxygen scaling parameter represents the attenuation due to the NH_3 overlayer. As discussed above, exposure of Al to O_2 under the present conditions produces thin Al_2O_3 (<7 Å) layers or islands even at exposures of 260 L. As seen in Table 2, the O_2 scaling factors decrease as the pre-oxidation exposure is increased. This indicates that the attenuation of the $O(2p)$ signal by the NH_3 overlayer is greater for the thicker oxide films. This is expected since the oxide thickness, even for a 260 L exposure, is

no more than twice the mean free path of the $O(2p)$ electron ($\lambda = 4 \text{ Å}$ [40,41]). Since a 0.9 L exposure of O_2 is completely incorporated just beneath the first layer of Al atoms without amorphous oxide formation, [45,46], attenuation of the oxygen UPS signal by a monolayer of NH_3 is not expected (NH_3 diameter = 3.04 Å). However, as thicker oxide layers (but still <7 Å thick) are formed by O_2 exposures of 4.3 and 260 L, NH_3 adsorption should cause attenuation of oxygen derived UPS intensities as observed.

From inspection of Fig. 9, one can see that the NH_3 ($1e^{-1}$) feature is shifted by 0.64 eV to higher BE when adsorbed on oxidized, as compared to clean, Al. A shift in this band could be due to a redistribution of charge in the N-H bonds as the result of hydrogen-bonding-like interactions with $O\delta^{-}$ anions (see below, however). It is difficult to tell directly from Fig. 9 whether the $(3a_1)^{-1}$ transition is shifted on oxidized Al. However, by subtracting the scaled $O(2p)$ resonance obtained on clean Al (dotted line), from the NH_3 covered oxidized Al surface (solid line), the resulting NH_3 features (not shown) indicate that the $(3a_1)^{-1}$ is also shifted by about 0.6 eV. The shift of both the $3a_1^{-1}$ and $1e^{-1}$ transitions by similar amounts suggests that, compared to clean Al, core hole screening is less effective in the oxidized case.

Campbell, et al. [58] found the same splitting of the $3a_1$ and $1e$ orbitals in solid ammonia as has been observed for the gas phase (see Table I). Referenced to the Fermi level of Al, the BE's of solid NH_3 are shifted 2.5 ± 0.1 eV to higher BE compared to those of NH_3 adsorbed on clean Al. A 0.64 eV shift to higher BE of the NH_3 levels following adsorption on oxidized Al, as well as the

absence of a significant differential shift between the $\tilde{\nu}$ 1e and $3a_1$ levels, suggests that the electronic environment of NH_3 in a solid molecular crystal and on an oxidized Al surface is similar. A shift of valence levels to higher BE has also been reported for CO adsorbed on oxygen-covered Pd [71]. Unfortunately, we can say nothing about how the peak widths change on oxidation due to uncertainties in the subtracted results.

The BE shift of the $\tilde{\nu}$ 1e transition is smaller in curve (a) where a low O_2 exposure (0.9 L) leaves little oxygen on the surface. The behavior of the resulting $Al\ \delta^+$ [45] surface to NH_3 adsorption should be similar to the clean Al surface except the bonding should be stronger due to the pre-existence of a permanent $Al\ \delta^+$ dipole rather than the weak dipole induced by NH_3 adsorption on clean Al. These results are comparable to NH_3 physisorption on the Zn-rich face of ZnO (0001) which exhibits the same splittings as gas phase NH_3 [26,72,73]. However, NH_3 chemisorbs on ZnO at 246 K, and accompanying this strong interaction, the splitting between the $\tilde{\nu}$ 1e⁻¹ and $\tilde{\nu}$ 3a⁻¹ transitions is 2 eV less than the gas phase. Comparison with physisorbed NH_3 shows that the $\tilde{\nu}$ 3a⁻¹ transition shifts with respect to the $\tilde{\nu}$ 1e⁻¹ indicating a strong lone-pair interaction with the surface. Although NH_3 adsorption on oxidized Al and Zn is similar in many respects the bond energy on ZnO is significantly stronger. The limiting work function changes, $\Delta\phi$, for saturation coverages of NH_3 and O_2 on clean Al are -1.7 and +0.2 eV respectively. Adsorption of NH_3 on an oxidized (260 L O_2 /58 L NH_3) surface leads to a maximum $\Delta\phi$ of -0.5 eV (with respect to clean Al). Clearly the magnitude of $\Delta\phi$ on the oxidized as compared

to the clean Al surface, even taking into account the small positive work function change which occurs upon oxidation of Al, indicates the NH_3 dipole is significantly different on the two surfaces. The sign of the work function change is still consistent with the NH_3 overlayer having the negative end of the NH_3 dipole pointing toward the surface. One possible explanation for the reduced magnitude of $\Delta\phi$ on oxidized compared to clean Al is that the hydrogen atoms on NH_3 are being stabilized by neighboring surface oxygen anions. This is consistent with the UPS results mentioned above. It would tend to decrease the magnitude of the NH_3 dipole and result in a decreased work function change (compared to $\Delta\phi_{max}$ observed for NH_3 on clean Al). However, we know of no good way to predict the work function change for this complex surface situation.

Work Function In light of the easy mobility of oxygen through the first few surface layers of Al, it is also possible that NH_3 adsorption induces changes in oxygen geometry and thereby work function changes unrelated to the NH_3 dipole.

In a separate set of experiments, the thermal stability of NH_3 adsorbed on an oxidized Al substrate was examined. The nitrogen-containing species were monitored using XPS of the N(1s) ionization. The sample was saturated with NH_3 at 128 K and then warmed slowly by terminating the flow of liquid nitrogen to the cooling coils. A two-step decrease in the area of the N(1s) peak at 400 eV with increasing temperature was observed indicating that roughly half the NH_3 desorbed at temperatures around 140 K. A small peak at 396 eV, assigned to adsorbed nitrogen atoms, was also observed.

its intensity increases during warmup to 200 K after which it declines. Characterization of the higher temperature species using thermal desorption and electron spectroscopic techniques warrants further investigation.

As noted earlier, XAES spectra of the N(KVV) region were taken for clean and oxidized Al surfaces saturated at 138 K with ammonia. The details are reported elsewhere [12]. By comparing the deconvoluted Auger spectra with gas phase [74,75, 76] and multilayer [77] spectra, perturbed NH_3 molecules are unmistakably identified as the major adsorbed species. In agreement with the UPS and XPS data reported here, the XAES data suggest that the $3a_1$ orbital is intimately involved in the interaction of NH_3 with the surface. The N(KVV) spectra for multilayer NH_3 and NH_3 adsorbed on oxidized Al are almost superimposable indicating the similarity in bonding in these two environments.

Although the bonding is similar, the thermal stability and the saturation coverage increase with the level of surface oxidation, from which we conclude that the bond strength increases as the Lewis acidity of the surface increases.

4. Summary and Conclusions.

In this work UPS, XPS, XAES and work function change measurements have been used to characterize the adsorption and bonding of NH_3 on clean and pre-oxidized Al surfaces. Where possible, comparisons are made between gas, monolayer and multilayer spectra. The following conclusions are drawn:

- (1) At 128 K, ammonia adsorbs molecularly on both clean and oxidized aluminum. There may be a small amount of dissociative adsorption.
- (2) The saturation coverage at 128 K lies between 25 and 50% of a close-packed NH_3 overlayer.
- (3) The initial sticking coefficient at 128 K is 0.13 ± 0.08 .
- (4) Multilayer formation is achieved by lowering the adsorption temperature to 106 K.
- (5) The thermal stability and extent of adsorption suggests that the NH_3 -substrate bond is stronger on oxidized than on clean Al. This may arise because there is a fairly strong dipole-dipole interaction on oxidized Al whereas on clean Al the interaction involves dipole-induced dipole forces.
- (6) Unlike NH_3 on polar ZnO and Fe where there is a 2-3 eV differential shift of $3a_1$ binding energies and unlike NH_3 ligands in organometallics where there is a strong dative bond, there is little shift of $3a_1$ and only a weak physisorption bond in the case of NH_3 on Al and oxidized Al.
- (7) The bonding of NH_3 to an oxidized Al surface and in a condensed molecular crystal is similar.

(8) The Lewis acidity of an oxidized Al surface increases as the extent of surface oxidation increases.

(9) There is no evidence for an NH_4^+ type species on clean or oxidized surfaces.

Two observations emerged in this study which might be applied to other systems. The first, which was only briefly mentioned in the text, is the use of condensates to enhance surface sensitivity. Interesting heterogeneous catalytic chemistry generally occurs on the surface, not beneath it. Electron spectroscopy offers excellent surface sensitivity compared to other techniques, but the signal may still originate many atomic layers below the surface; the observed intensity is thus a combination of surface and bulk transitions. Preferential attenuation of the bulk signal by an inert, easily vaporized overlayer offers a powerful technique for determining, in a non-destructive fashion, interactions occurring within the first few atomic layers.

Secondly, UPS superposition techniques have been successful in separating intrinsic adsorbate effects from those extrinsically induced by the surface and vice versa. Sensible and useful information has been gained by interpreting the significance of scaling parameters used in the summation of spectra.

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TABLE II
Scaling Factors for Ammonia on Pre-oxidized Aluminum

	Oxygen Pre-dose		
	0.9 L	4.3 L	260 L
Scaling factor for oxygen features ^a	1.0	0.8	0.7
Scaling factor for NH ₃ features ^a	1.1	1.1	1.4
Ratio of N(1s)[oxidized]/N(1s)[clean] ^b	1.8	1.8	2.3

^afrom Fig. 9

^bfrom Fig. 8

^aAll energies are reported in eV.
^bSaturation coverage (115 L exposure) at 128 K.
^cThe work function at saturation NH₃ coverage has been subtracted from the reported literature values to facilitate comparison with the present work ($\Phi_{A1} + \Delta\Phi_{sat} = 2.5$ eV).
^dThe instrument resolution was not reported in Ref. 58 but we assume that it was at least better than $\Delta E(FWHM) = 0.4$ eV since the work was done in the gas phase.

Spectrum	Gaseous NH ₃			Solid NH ₃			Adsorbed NH ₃	
	3a ₁	1e	1a ₁	3a ₁	1e	3a ₁	3a ₁	1a ₁
Binding Energies ^c	8.38 [49]	13.5 [49]	403.5 [51]	8.8 [58]	14.2 [58]	6.3	11.8	400.4
Band Widths	1.7 ^d [57]	2.5 [57]	--	1.8 [57]	3.0 [58]	2.6	3.7	--
Energy Separation	5.12 [49]			5.4 [58]			5.5	
Branching Ratio 1e/3a ₁	2.0 [54]			2.18 [58]			2.29	
--	--			--			--	

TABLE I
Electron Spectroscopic Parameters for Ammonia

Figure Captions.

Figure 1. O(1s) spectra for various exposures of polycrystalline Al to O₂ at 128 K. Mg K α x-rays at 1253.6 eV.

Figure 2. He I and He II photoelectron spectra for oxygen chemisorbed on polycrystalline Al at 128 K. Note different intensity scales for the two panels.

Figure 3. Peak areas for oxygen valence (O) by UPS and core 1s (Δ) by XPS photoionizations as a function of oxygen exposure at 128 K. Solid line is a fit to a first order adsorption model.

Figure 4. N(1s) core level XPS spectra for various exposures of NH₃ on polycrystalline Al. Spectra a-e involve exposure at 128 K. Spectrum f is a multilayer built up at 106 K.

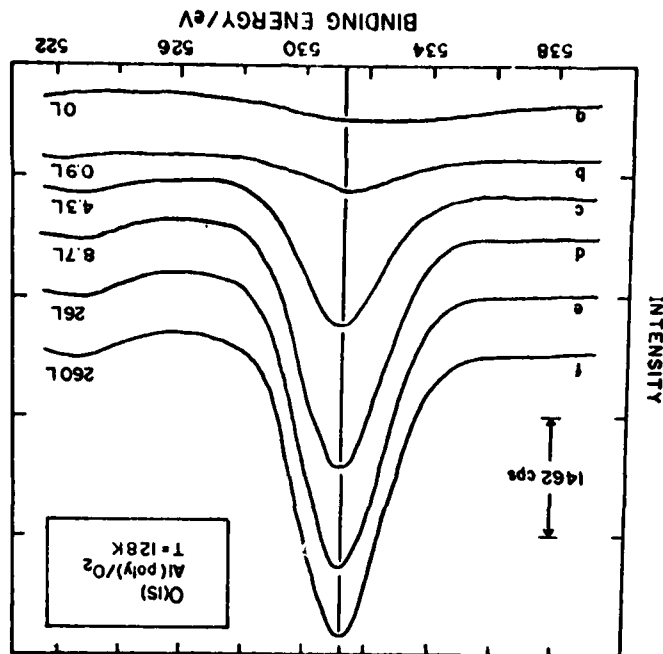
Figure 5. Gas phase UPS/Ammonia [49] and representations of valence molecular orbitals of NH₃ showing contours of constant charge density. (Reproduced from reference 48 with permission.)

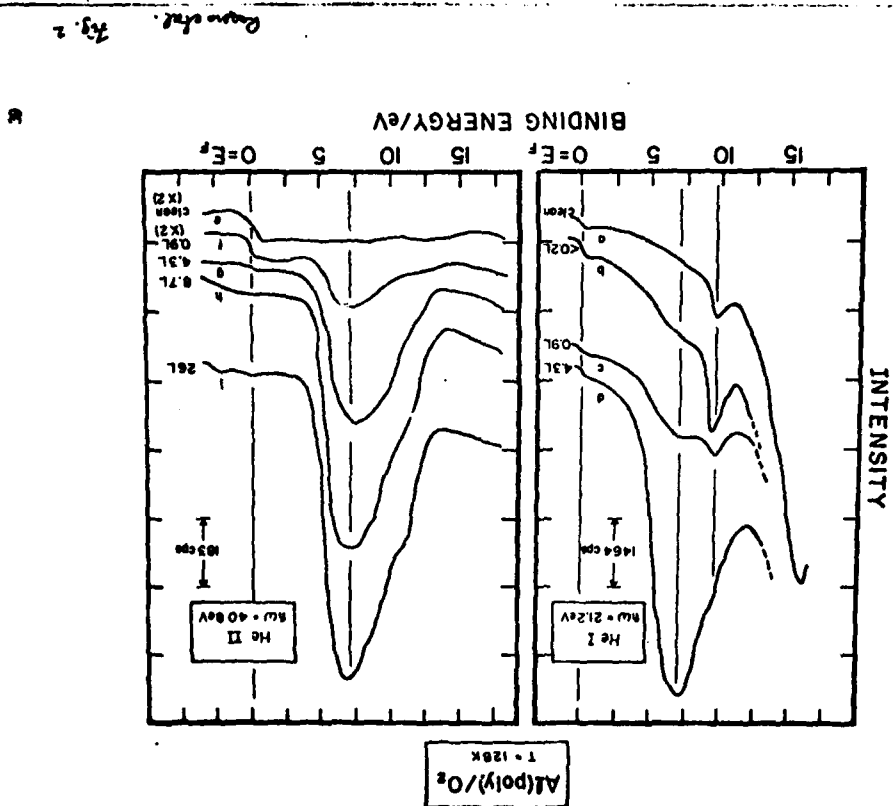
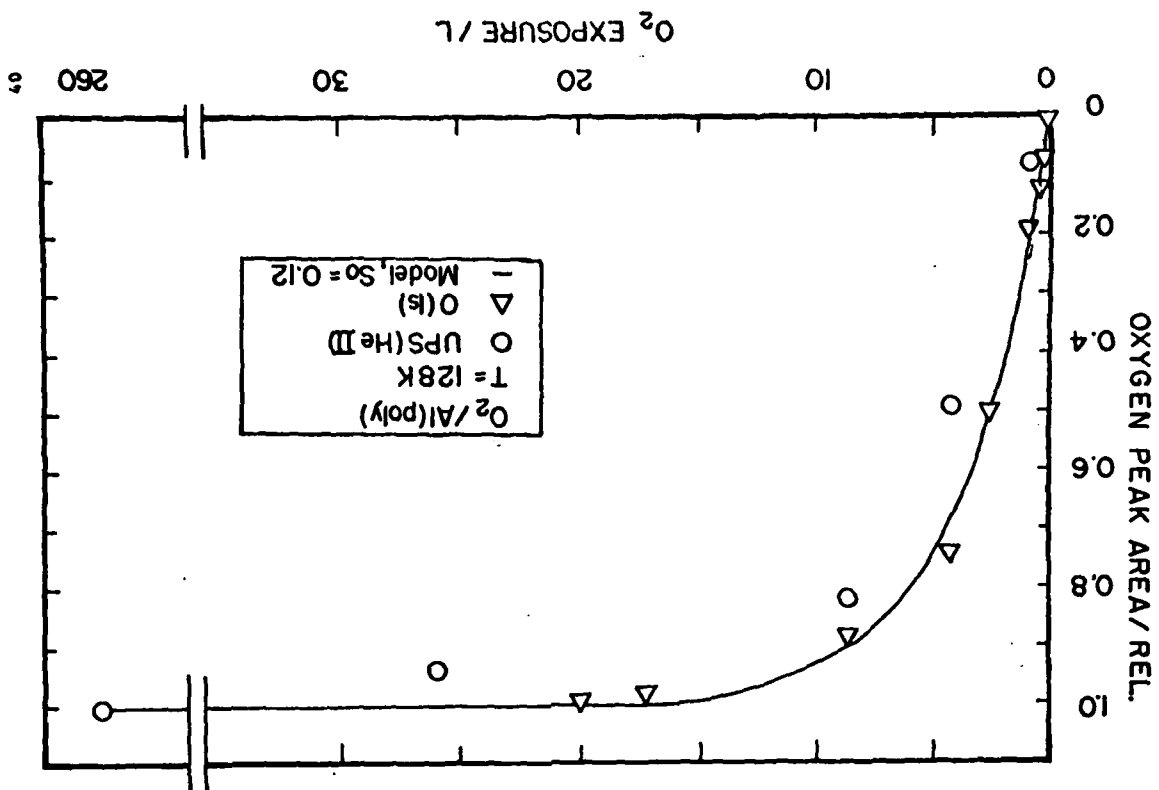
Figure 6. UPS (He II) spectra for various exposures of NH₃ on polycrystalline Al at 128 K. The gas phase ionization potentials (see text) are shown at the top.

Figure 7. Peak areas as a function of exposure at 128 K for the x-ray excited N(1s) core level (Δ) and the He II valence levels (O) of ammonia adsorbed on polycrystalline Al. The solid line is a fit to a first order Langmuir adsorption model with an initial sticking coefficient of 0.13 (see text). The inset shows the change in work function with ammonia exposure. This was determined using the widths of He II spectra.

Figure 8. N(1s) peak areas, normalized to the Al (2s) peak area for a surface free of ammonia, versus exposure at 128 K (O) and 106 K (Δ). The filled squares show the same quantity, determined at 128 K, for Al preexposed in O₂ at three different exposures followed by a 58 L exposure to NH₃.

Figure 9. He II spectra for NH₃ adsorbed on oxidized Al at 128 K. Solid curve is actual UPS spectrum. Dotted curve is a scaled spectrum of oxygen alone on Al. Dashed curve is a scaled version of NH₃ alone on Al. Dot-dashed curve is sum of the two scaled spectra. A involves a preexposure of 0.9 L O₂ followed by 58 L NH₃. B is for 4.3 L O₂ and 58 L NH₃ while C is for 260 L O₂ and 58 L NH₃.





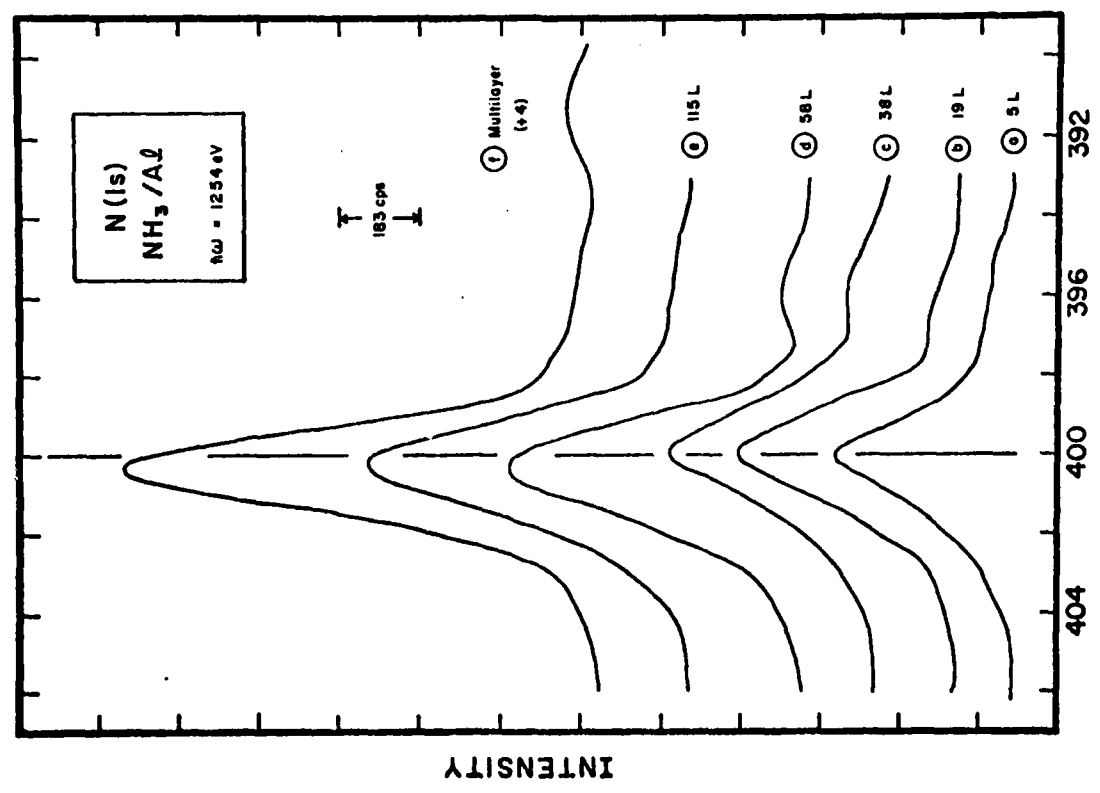


Fig. 4
Rynn et al.

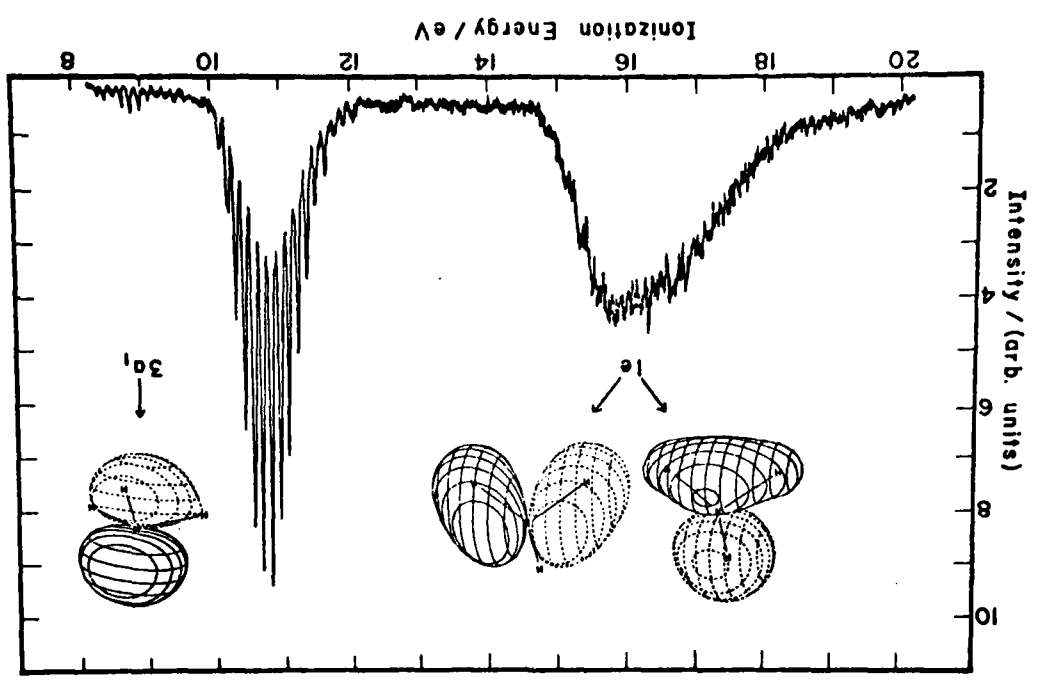


Fig. 5
Rynn et al.

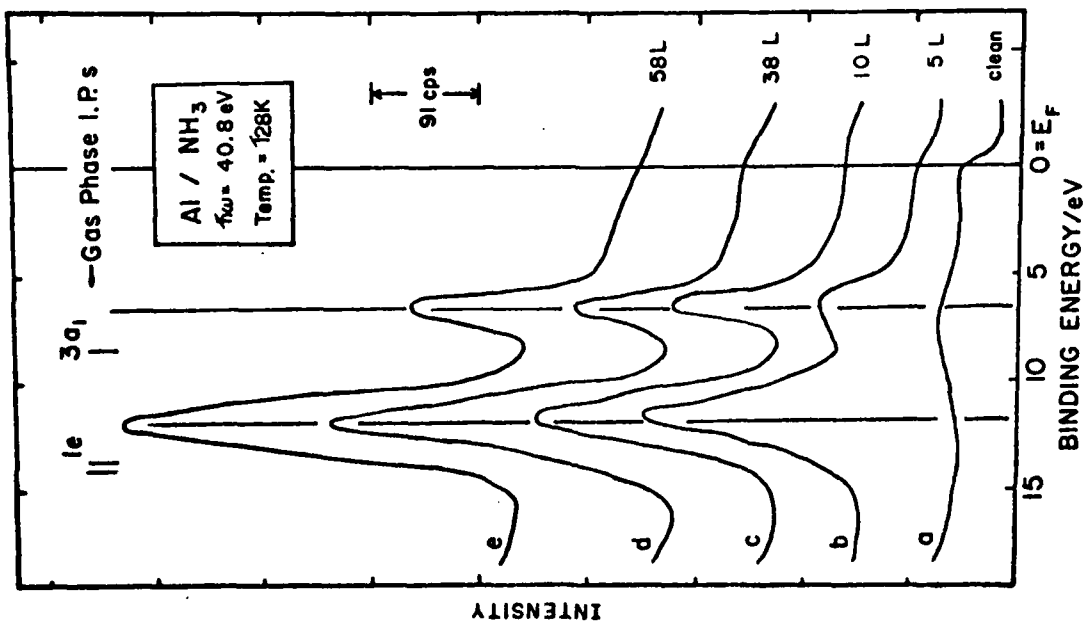
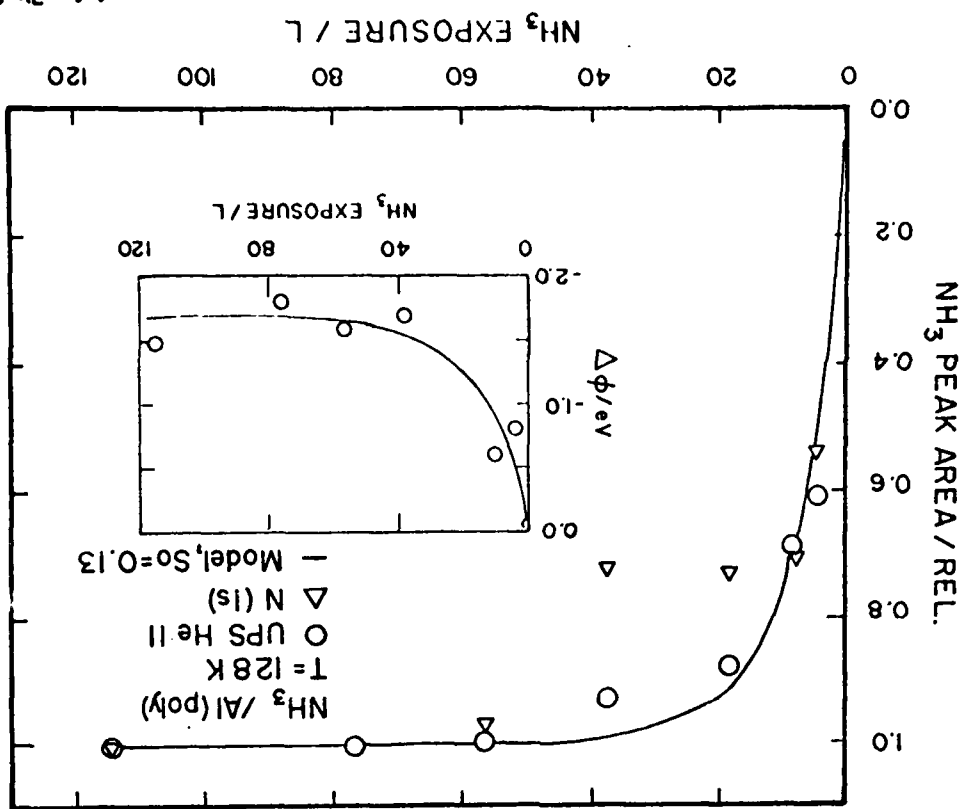


Fig. 6



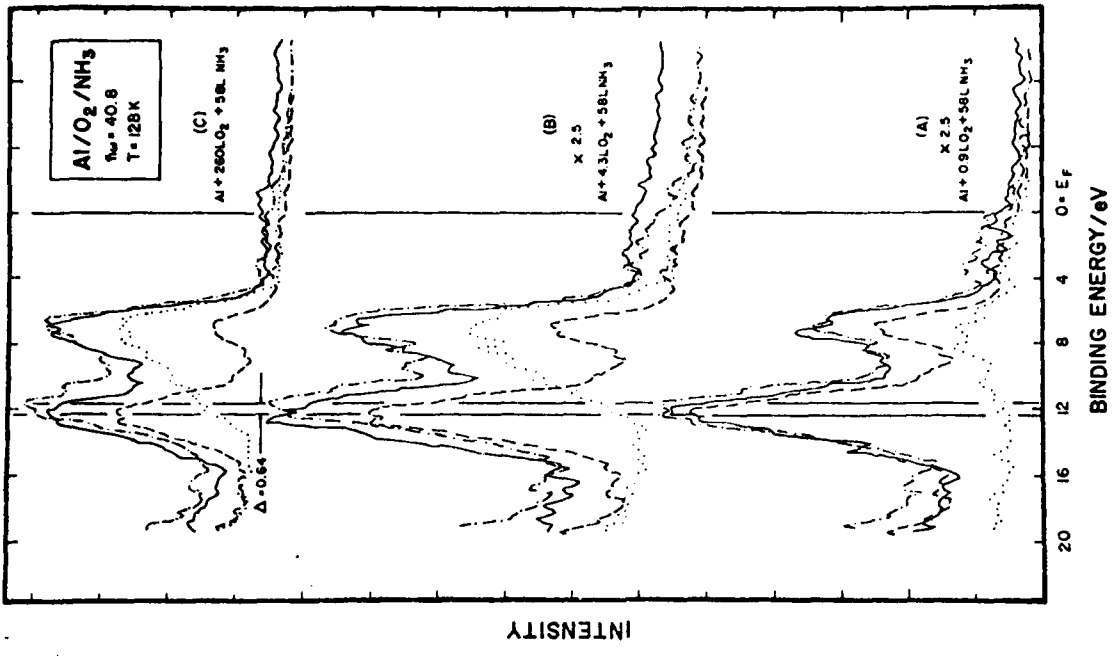


Figure 46. XPS spectra of Al 2p for Al/O₂/NH₃ system at 128 K.

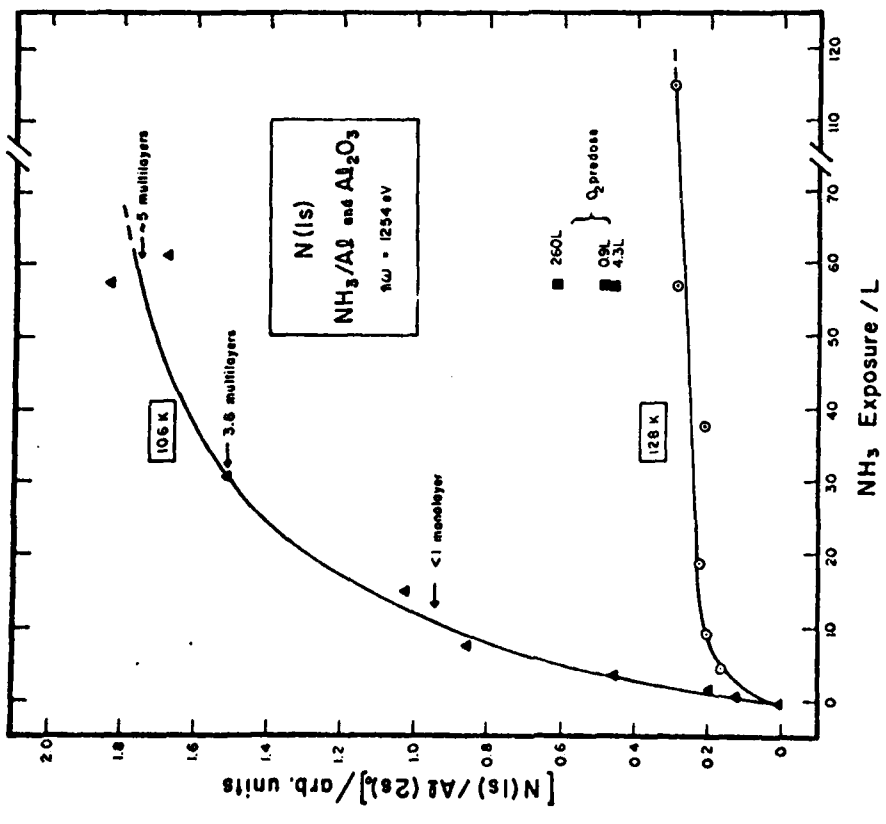


Figure 45. Plot of N(1s)/Al(2s) vs NH₃ Exposure for Al₂O₃ at 106 K and 128 K.